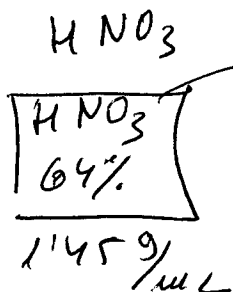


(22)



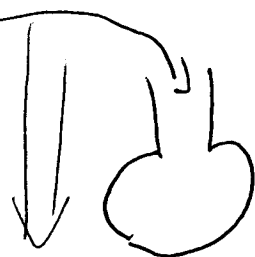
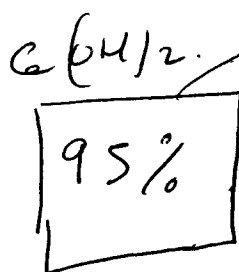
250 mL } 0'375 mols \Rightarrow
 1'5 M.

$\Rightarrow 23'625 \text{ g S} = 0'375 \cdot 63 \text{ g/mol}$

$23'625 \text{ g S} \cdot \frac{100 \text{ g D}}{64 \text{ g S}} = 36'9 \text{ g D}$

$V_D = \frac{36'9}{1'45} = 25'46 \text{ mL}$

(23)



100 mL } 0'08 mols Soluto
 0'8 M. $\Rightarrow 0'08 \cdot 74 \text{ g} = 5'92 \text{ g S}$

$5'92 \text{ g Soluto} \cdot \frac{95 \text{ g S}}{100 \text{ g D}} = 6'2 \text{ g D}$

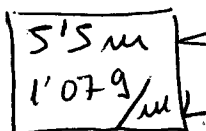
(24)



100 mL } 0'02 mols Soluto
 0'2 M

$V_D = \frac{0'02}{0'5} = 0'04 \text{ L} = 40 \text{ mL}$

(26)



5'5 mols
 1000 g d.

$M = \frac{5'5 \text{ mols}}{1'122 \text{ L}} = 4'9 \text{ M}$

$M_D = 1000 + 5'5 \cdot 36'5 = 1200'75 \text{ g D}$

$P_m = 36'5 \text{ g/mol}$

$V_D = \frac{1200'75}{1'07} = 1122'2 \text{ mL}$

$X_S = \frac{5'5}{5'5 + \frac{1000}{18}} = 0'09$

$\% = \frac{5'5 \cdot 36'5}{1200'75} \cdot 100 = 16'7\%$